**3. PARTICLE MODEL OF MATTER**

1. Density
2. *Practical: Investigate the densities of regular and irregular solid objects and liquids*
3. Changes of state
4. Internal energy
5. *Perform an experiment to measure the latent heat of fusion of water.*
6. Specific heat capacity
7. Latent heat
8. Particle motion in gases
9. Increasing the pressure of a gas
10. Particle model and change of state

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| **4.3 Particle model of matter** |  |  |  |
| *4.3.1 Changes of state and the particle model* |  |  |  |
| **4.3.1.1 Density of materials** |  |  |  |
| a) The density of a material is defined by the equation:  density, *ρ*, in kilograms per metre cubed, kg/m**3**  mass, *m*, in kilograms, kg  volume, *V*, in metres cubed, m**3**  b) The particle model can be used to explain   * the different states of matter * differences in density.   c) Recognise/draw simple diagrams to model the difference between solids, liquids and gases.  d) Explain the differences in density between the different states of matter in terms of the arrangement of atoms or molecules. |  |  |  |
| **Required practical activity 5:** use appropriate apparatus to make and record the measurements needed to determine the densities of regular and irregular solid objects and liquids. Volume should be determined from the dimensions of regularly shaped objects, and by a displacement technique for irregularly shaped objects. Dimensions to be measured using appropriate apparatus such as a ruler, micrometer or Vernier callipers. |  |  |  |
| **4.3.1.2 Changes of state** |  |  |  |
| 1. Describe how, when substances change state (melt, freeze, boil, evaporate, condense or sublimate), mass is conserved. 2. Changes of state are physical changes which differ from chemical changes because the material recovers its original properties if the change is reversed. |  |  |  |
| *4.3.2 Internal energy and energy transfers* |  |  |  |
| **4.3.2.1 Internal energy** |  |  |  |
| 1. Energy is stored inside a system by the particles (atoms and molecules) that make up the system. This is called internal energy. 2. Internal energy is the total kinetic energy and potential energy of all the particles (atoms and molecules) that make up a system. 3. Heating changes the energy stored within the system by increasing the energy of the particles that make up the system. This either raises the temperature of the system or produces a change of state. |  |  |  |
| **4.3.2.2 Temperature changes in a system and specific heat capacity** |  |  |  |
| a) If the temperature of the system increases, the increase in temperature depends on the mass of the substance heated, the type of material and the energy input to the system.  b) The following equation applies:  change in thermal energy, Δ*E*, in joules, J  mass, *m*, in kilograms, kg  specific heat capacity, *c*, in joules per kilogram per degree Celsius, J/kg °C  temperature change, Δ*θ*, in degrees Celsius, °C.  c) The specific heat capacity of a substance is the amount of energy required to raise the temperature of one kilogram of the substance by one degree Celsius. |  |  |  |
| **4.3.2.3 Changes of heat and specific latent heat** |  |  |  |
| If a change of state happens:   1. The energy needed for a substance to change state is called latent heat. 2. When a change of state occurs, the energy supplied changes the energy stored (internal energy) but not the temperature. 3. The specific latent heat of a substance is the amount of energy required to change the state of one kilogram of the substance with no change in temperature.   energy, *E*, in joules, J  mass, *m*, in kilograms, kg  specific latent heat, *L*, in joules per kilogram, J/kg  Specific latent heat of *fusion* is the change of state from solid to liquid  Specific latent heat of *vaporisation* is the change of state from liquid to vapour  **Perform an experiment** to measure the latent heat of fusion of water.  Be able to interpret heating and cooling graphs that include changes of state.  Be able to distinguish between specific heat capacity and specific latent heat. |  |  |  |
| *4.3.3 Particle model and pressure* |  |  |  |
| **4.3.3.1 Particle motion in gases** |  |  |  |
| 1. The molecules of a gas are in constant random motion. 2. The temperature of the gas is related to the average kinetic energy of the molecules. 3. Changing the temperature of a gas, held at constant volume, changes the pressure exerted by the gas 4. Be able to:  * explain how the motion of the molecules in a gas is related to both its temperature and its pressure * explain qualitatively the relation between the temperature of a gas and its pressure at constant volume. |  |  |  |
| **4.3.3.2 Pressure in gases (physics only)** |  |  |  |
| 1. A gas can be compressed or expanded by pressure changes. 2. The pressure produces a net force at right angles to the wall of the gas container (or any surface). 3. Be able to use the particle model to explain how increasing the volume in which a gas is contained, at constant temperature, can lead to a decrease in pressure. 4. For a fixed mass of gas held at a constant temperature:   pressure, *p*, in pascals, Pa  volume, *V*, in metres cubed, m**3**  e) Calculate the change in the pressure of a gas or the volume of a gas (a fixed mass held at constant temperature) when either the pressure or volume is increased or decreased. |  |  |  |
| **4.3.3.3 Increasing the pressure of a gas (physics only) (HT only)** |  |  |  |
| 1. Work is the transfer of energy by a force. 2. Doing work on a gas increases the internal energy of the gas and can cause an increase in the temperature of the gas. 3. Explain how, in a given situation eg a bicycle pump, doing work on an enclosed gas leads to an increase in the temperature of the gas. |  |  |  |

**PHYSICS EQUATIONS TO LEARN BY HEART**

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| **PARTICLE MODEL topic** | | | | | | |
|  | | **Quantity** | | **Unit** | | **Equation** |
| *PART-ICLES* | *13* | **𝞀**  **m**  **v** | density  mass  volume | kilogram per metre3  kilogram  metre3 | kg/m3  kg  m3 |  |